

Name: _____

Ionic Compounds Review

Directions: Use the periodic table to find the charges of the ions and write them into the correct ionic compound. Then write the name of the compound

Formula Writing Rules:

1. Ionic compounds are NEUTRAL so the total charge of the positive ions must equal the total charge of the negative ions.
2. Write the formula of the positive ion (cation) first followed by the negative ion (anion).
3. Put parenthesis around polyatomic ions if (and only if) there are more than one of these ions in a compound.

Naming Rules:

- 1) Write the name of the first element
- 2) Write the name of the second element with an “-ide” ending.
- 3) ***If the compound contains a polyatomic ion, simply write the name for that polyatomic ion without changing the name.
 - a. Write the name of the element followed by the name for the polyatomic ion.
 - b. If it starts with NH_4 (ammonium) write ammonium followed by:
 - i. the name of the following element with the -ide ending or
 - ii. the name of the following polyatomic ion.

	Elements		Cation	Anion	Ionic Compound Formula	Compound Name
1	Ba	O				
2	Mg	Cl				
3	Rb	S				
4	Al	Br				
5	Na	Cl				
6	Ca	N				
7	Ba	N				
8	Ca	Cl				
9	Al	O				
10	B	Br				

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11	Sn	Se				
12	Ga	As				
13	Mg	Br				
14	Li	I				
15	K	O				
16	Mg	S				
17	Na	F				
18	Al	Cl				
19	K	SO ₄				
20	NH ₄	Se				
21	Sr	NO ₃				
22	K	PO ₄				
23	NH ₄	Cl				
24	Ba	OH				
25	Li	SO ₄				
26	NH ₄	PO ₄				
27	Be	HCO ₃				
28	K	CN				
29	Sr	OH				
30	Al	PO ₄				
31	Na	MnO ₄				
32	NH ₄	HCO ₃				
33	Sn	NO ₂				
34	Al	CN				
35	Ca	NO ₂				