

Naming Ionic Compounds Notes

Step 1: Decide what type of compound it is.

- An ionic compound is made up of a metal and a nonmetal.

- 1) Write the name of the first element
- 2) Write the name of the second element with an “-ide” ending.

<u>Elements</u>	<u>Ending with “-ide”</u>
Oxygen	oxide
Nitrogen	nitride
Sulfur	sulfide
Chlorine	chloride
Fluorine	fluoride
Bromine	bromide

Ex: $\text{KCl} =$ Potassium chloride

$\text{NaCl} =$ Sodium chloride

***If the compound contains a polyatomic ion:

- a. Write the name of the element followed by the name for the polyatomic ion.

i. Ex: $\text{Na (OH)} =$ Sodium hydroxide

- b. If it starts with NH_4 (ammonium) write ammonium followed by either:

- i. the name of the following element with the -ide ending

1. Ex: $(\text{NH}_4)_2\text{O} =$ Ammonium Oxide

- ii. OR the name of the following polyatomic ion.

1. $\text{NH}_4\text{OH} =$ Ammonium Hydroxide